

BEST AVAILABLE COPY

Table 1 (continued)
ALPHABETICAL LISTING

Reaction	E°, V	Reaction	E°, V
$\text{PbSO}_4 + 2e \rightleftharpoons \text{Pb(Hg)} + \text{SO}_4^{2-}$	-0.3505	$\text{Se} + 2\text{H}^+ + 2e \rightleftharpoons \text{H}_2\text{Se(aq)}$	-0.399
$\text{Pd}^{2+} + 2e \rightleftharpoons \text{Pd}$	0.951	$\text{H}_2\text{SeO}_3 + 4\text{H}^+ + 4e \rightleftharpoons \text{Se} + 3\text{H}_2\text{O}$	-0.74
$[\text{PdCl}_4]^{2-} + 2e \rightleftharpoons \text{Pd} + 4\text{Cl}^-$	0.591	$\text{SeO}_3^{2-} + 3\text{H}_2\text{O} + 4e \rightleftharpoons \text{Se} + 6\text{OH}^-$	-0.365
$[\text{PdCl}_6]^{4-} + 2e \rightleftharpoons [\text{PdCl}_4]^{2-} + 2\text{Cl}^-$	-1.288	$\text{SeO}_2^{2-} + 4\text{H}^+ + 2e \rightleftharpoons \text{H}_2\text{SeO}_3 + \text{H}_2\text{O}$	1.151
$\text{Pd(OH)}_2 + 2e \rightleftharpoons \text{Pd} + 2\text{OH}^-$	0.07	$\text{SeO}_2 + \text{H}_2\text{O} + 2e \rightleftharpoons \text{SeO}_3^{2-} + 2\text{OH}^-$	0.05
$\text{Pt}^{2+} + 2e \rightleftharpoons \text{Pt}$	-1.118	$\text{SiF}_6^{2-} + 4e \rightleftharpoons \text{Si} + 6\text{F}^-$	-1.24
$[\text{PtCl}_4]^{2-} + 2e \rightleftharpoons \text{Pt} + 4\text{Cl}^-$	0.755	$\text{SiO}_2 \text{ (quartz)} + 4\text{H}^+ + 4e \rightleftharpoons \text{Si} + 2\text{H}_2\text{O}$	-0.857
$[\text{PtCl}_6]^{4-} + 2e \rightleftharpoons [\text{PtCl}_4]^{2-} + 2\text{Cl}^-$	0.68	$\text{SiO}_3^{2-} + 3\text{H}_2\text{O} + 4e \rightleftharpoons \text{Si} + 6\text{OH}^-$	-1.697
$\text{Pt(OH)}_2 + 2e \rightleftharpoons \text{Pt} + 2\text{OH}^-$	0.14	$\text{Sn}^{2+} + 2e \rightleftharpoons \text{Sn}$	-0.1375
$\text{Pu}^{3+} + 3e \rightleftharpoons \text{Pu}$	-2.031	$\text{Sn}^{4+} + 2e \rightleftharpoons \text{Sn}^{2+}$	0.151
$\text{Pu}^{4+} + e \rightleftharpoons \text{Pu}^{3+}$	1.006	$\text{HSnO}_2^- + \text{H}_2\text{O} + 2e \rightleftharpoons \text{Sn} + 3\text{OH}^-$	-0.909
$\text{Pu}^{4+} + e \rightleftharpoons \text{Pu}^{3+}$	1.099	$\text{HSnO}_2^- + \text{H}_2\text{O} + 2e \rightleftharpoons \text{Sn} + 3\text{OH}^-$	-0.93
$\text{PuO}_2(\text{OH})_2 + 2\text{H}^+ + 2e \rightleftharpoons \text{Pu(OH)}_4$	1.325	$\text{Sn(OH)}_6^{2-} + 2e \rightleftharpoons \text{HSnO}_2^- + 3\text{OH}^- + \text{H}_2\text{O}$	-4.10
$\text{PuO}_2(\text{OH})_2 + \text{H}^+ + e \rightleftharpoons \text{PuO}_2\text{OH} + \text{H}_2\text{O}$	-1.062	$\text{Sr}^{2+} + 2e \rightleftharpoons \text{Sr}$	-2.89
$\text{Rb}^+ + e \rightleftharpoons \text{Rb}$	0.300	$\text{Sr}^{2+} + 2e \rightleftharpoons \text{Sr(Hg)}$	-2.88
$\text{Re}^{3+} + 3e \rightleftharpoons \text{Re}$	0.510	$\text{Sr(OH)}_2 + 2e \rightleftharpoons \text{Sr} + 2\text{OH}^-$	-0.750
$\text{ReO}_4^- + 4\text{H}^+ + 3e \rightleftharpoons \text{ReO}_2 + 2\text{H}_2\text{O}$	0.2513	$\text{Ta}_2\text{O}_5 + 10\text{H}^+ + 10e \rightleftharpoons 2\text{Ta} + 5\text{H}_2\text{O}$	0.400
$\text{ReO}_2 + 4\text{H}^+ + 4e \rightleftharpoons \text{Re} + 2\text{H}_2\text{O}$	0.768	$\text{Ta}_2\text{O}_5 + 10\text{H}^+ + 10e \rightleftharpoons 2\text{Ta} + 5\text{H}_2\text{O}$	0.782
$\text{ReO}_3 + 2\text{H}^+ + e \rightleftharpoons \text{ReO}_4^- + \text{H}_2\text{O}$	-0.584	$\text{Tc}^{2+} + 2e \rightleftharpoons \text{Tc}$	0.782
$\text{ReO}_4^- + 4\text{H}_2\text{O} + 7e \rightleftharpoons \text{Re} + 8\text{OH}^-$	0.368	$\text{TcO}_4^- + 4\text{H}^+ + 3e \rightleftharpoons \text{TcO}_2 + 2\text{H}_2\text{O}$	-1.143
$\text{ReO}_4^- + 8\text{H}^+ + 7e \rightleftharpoons \text{Re} + 4\text{H}_2\text{O}$	0.600	$\text{Te} + 2e \rightleftharpoons \text{Te}^{2-}$	-0.793
$\text{Rh}^+ + e \rightleftharpoons \text{Rh}$	0.600	$\text{Te} + 2\text{H}^+ + 2e \rightleftharpoons \text{H}_2\text{Te}$	0.568
$\text{Rh}^{2+} + 2e \rightleftharpoons \text{Rh}$	-0.758	$\text{Te}^{4+} + 4e \rightleftharpoons \text{Te}$	-0.593
$\text{Rh}^{3+} + 3e \rightleftharpoons \text{Rh}$	0.431	$\text{TeO}_2 + 4\text{H}^+ + 4e \rightleftharpoons \text{Te} + 2\text{H}_2\text{O}$	-0.57
$[\text{RhCl}_6]^{3-} + 3e \rightleftharpoons \text{Rh} + 6\text{Cl}^-$	0.455	$\text{TeO}_2 + 4\text{H}^+ + 4e \rightleftharpoons \text{Te} + 2\text{H}_2\text{O}$	-0.472
$\text{Ru}^{2+} + 2e \rightleftharpoons \text{Ru}$	0.2487	$\text{TeO}_3^{2-} + 3\text{H}_2\text{O} + 4e \rightleftharpoons \text{Te} + 6\text{OH}^-$	1.02
$\text{Ru}^{3+} + e \rightleftharpoons \text{Ru}^{2+}$	1.120	$\text{TeO}_4^{2-} + 8\text{H}^+ + 7e \rightleftharpoons \text{Te} + 4\text{H}_2\text{O}$	-1.899
$\text{RuO}_2 + 4\text{H}^+ + 2e \rightleftharpoons \text{Ru}^{2+} + 2\text{H}_2\text{O}$	0.59	$\text{H}_6\text{TeO}_6 + 2\text{H}^+ + 2e \rightleftharpoons \text{TeO}_2 + 4\text{H}_2\text{O}$	-1.789
$\text{RuO}_4^- + e \rightleftharpoons \text{RuO}_4^{2-}$	1.00	$\text{Th}^{4+} + 4e \rightleftharpoons \text{Th}$	-2.48
$\text{RuO}_4 + e \rightleftharpoons \text{RuO}_4^-$	-0.47627	$\text{ThO}_2 + 4\text{H}^+ + 4e \rightleftharpoons \text{Th} + 2\text{H}_2\text{O}$	-1.630
$\text{S} + 2e \rightleftharpoons \text{S}^{2-}$	0.142	$\text{Th(OH)}_4 + 4e \rightleftharpoons \text{Th} + 4\text{OH}^-$	-0.368
$\text{S} + 2\text{H}^+ + 2e \rightleftharpoons \text{H}_2\text{S(aq)}$	-0.478	$\text{Ti}^{2+} + 2e \rightleftharpoons \text{Ti}$	-0.502
$\text{S} + \text{H}_2\text{O} + 2e \rightleftharpoons \text{HS}^- + \text{OH}^-$	-0.42836	$\text{Ti}^{3+} + e \rightleftharpoons \text{Ti}^{2+}$	-0.055
$2\text{S} + 2e \rightleftharpoons \text{S}_2^{2-}$	0.564	$\text{TiO}_2 + 4\text{H}^+ + 2e \rightleftharpoons \text{Ti}^{2+} + 2\text{H}_2\text{O}$	-0.336
$\text{S}_2\text{O}_8^{2-} + 4\text{H}^+ + 2e \rightleftharpoons 2\text{H}_2\text{SO}_4$	2.010	$\text{TiOH}^{3+} + \text{H}^+ + e \rightleftharpoons \text{Ti}^{3+} + \text{H}_2\text{O}$	-0.3338
$\text{S}_2\text{O}_8^{2-} + 2e \rightleftharpoons 2\text{SO}_4^{2-}$	2.123	$\text{Ti}^+ + e \rightleftharpoons \text{Ti(Hg)}$	1.252
$\text{S}_2\text{O}_8^{2-} + 2\text{H}^+ + 2e \rightleftharpoons 2\text{HSO}_4^-$	0.08	$\text{Ti}^+ + e \rightleftharpoons \text{Ti}$	-0.658
$\text{S}_4\text{O}_8^{2-} + 2e \rightleftharpoons 2\text{S}_2\text{O}_8^{2-}$	-0.056	$\text{TiBr} + e \rightleftharpoons \text{Ti} + \text{Br}^-$	-0.5568
$2\text{H}_2\text{SO}_3 + \text{H}^+ + 2e \rightleftharpoons \text{HS}_2\text{O}_4^- + 2\text{H}_2\text{O}$	0.449	$\text{TiCl} + e \rightleftharpoons \text{Ti} + \text{Cl}^-$	-0.752
$\text{H}_2\text{SO}_3 + 4\text{H}^+ + 4e \rightleftharpoons \text{S} + 3\text{H}_2\text{O}$	-1.12	$\text{TiI} + e \rightleftharpoons \text{Ti} + \text{I}^-$	0.02
$2\text{SO}_3^{2-} + 2\text{H}_2\text{O} + 2e \rightleftharpoons \text{S}_2\text{O}_4^{2-} + 4\text{OH}^-$	-0.571	$\text{TiI}_2 + e \rightleftharpoons \text{Ti} + \text{I}^-$	-0.34
$2\text{SO}_3^{2-} + 3\text{H}_2\text{O} + 4e \rightleftharpoons \text{S}_2\text{O}_3^{2-} + 6\text{OH}^-$	0.172	$\text{TiO}_3 + 3\text{H}_2\text{O} + 4e \rightleftharpoons 2\text{Ti}^{2+} + 6\text{OH}^-$	-0.05
$\text{SO}_4^{2-} + 4\text{H}^+ + 2e \rightleftharpoons \text{H}_2\text{SO}_3 + \text{H}_2\text{O}$	-0.22	$\text{TiOH} + e \rightleftharpoons \text{Ti} + \text{OH}^-$	-0.4360
$2\text{SO}_4^{2-} + 4\text{H}^+ + 2e \rightleftharpoons \text{S}_2\text{O}_8^{2-} + \text{H}_2\text{O}$	-0.93	$\text{Ti(OH)}_3 + 2e \rightleftharpoons \text{TiOH} + 2\text{OH}^-$	-1.798
$\text{SO}_4^{2-} + \text{H}_2\text{O} + 2e \rightleftharpoons \text{SO}_3^{2-} + 2\text{OH}^-$	-0.510	$\text{Ti}_2\text{SO}_4 + 2e \rightleftharpoons \text{Ti} + \text{SO}_4^{2-}$	-0.607
$\text{Sb} + 3\text{H}^+ + 3e \rightleftharpoons \text{SbH}_3$	0.152	$\text{U}^{3+} + 3e \rightleftharpoons \text{U}$	0.612
$\text{Sb}_2\text{O}_3 + 6\text{H}^+ + 6e \rightleftharpoons 2\text{Sb} + 3\text{H}_2\text{O}$	0.671	$\text{U}^{4+} + e \rightleftharpoons \text{U}^{3+}$	0.062
$\text{Sb}_2\text{O}_3 \text{ (senarmonite)} + 4\text{H}^+ + 4e \rightleftharpoons \text{Sb}_2\text{O}_3 + 2\text{H}_2\text{O}$	0.649	$\text{UO}_2^{2+} + 4\text{H}^+ + e \rightleftharpoons \text{U}^{4+} + 2\text{H}_2\text{O}$	0.327
$\text{Sb}_2\text{O}_3 \text{ (valentinite)} + 4\text{H}^+ + 4e \rightleftharpoons \text{Sb}_2\text{O}_3 + 2\text{H}_2\text{O}$	0.581	$\text{UO}_3^{2+} + e \rightleftharpoons \text{UO}^{2+}$	-1.444
$\text{Sb}_2\text{O}_3 + 6\text{H}^+ + 4e \rightleftharpoons 2\text{SbO}^+ + 3\text{H}_2\text{O}$	0.212	$\text{UO}_3^{2+} + 4\text{H}^+ + 2e \rightleftharpoons \text{U}^{4+} + 2\text{H}_2\text{O}$	-1.175
$\text{SbO}^+ + 2\text{H}^+ + 3e \rightleftharpoons \text{Sb} + 2\text{H}_2\text{O}$	-0.66	$\text{UO}_3^{2+} + 4\text{H}^+ + 6e \rightleftharpoons \text{U} + 2\text{H}_2\text{O}$	-0.255
$\text{SbO}_2^- + 2\text{H}_2\text{O} + 3e \rightleftharpoons \text{Sb} + 4\text{OH}^-$	-0.59	$\text{V}^{2+} + 2e \rightleftharpoons \text{V}$	0.337
$\text{SbO}_2^- + 2\text{H}_2\text{O} + 3e \rightleftharpoons \text{SbO}_2^- + 2\text{OH}^-$	-2.077	$\text{V}^{3+} + e \rightleftharpoons \text{V}^{2+}$	0.991
		$\text{VO}^{2+} + 2\text{H}^+ + e \rightleftharpoons \text{V}^{3+} + \text{H}_2\text{O}$	1.00
		$\text{VO}_2^+ + 2\text{H}^+ + e \rightleftharpoons \text{VO}^{2+} + \text{H}_2\text{O}$	-0.254
		$\text{V(OH)}_4^+ + 2\text{H}^+ + e \rightleftharpoons \text{VO}^{2+} + 3\text{H}_2\text{O}$	-0.031
		$\text{V(OH)}_4^+ + 4\text{H}^+ + 5e \rightleftharpoons \text{V} + 4\text{H}_2\text{O}$	
		$\text{w.O.} + 2\text{H}^+ + 2e \rightleftharpoons 2\text{WO}_2 + \text{H}_2\text{O}$	

WO₂ + 4H⁺ + 4e⁻ =
WO₃ + 6H⁺ + 6e⁻ =
2WO₃ + 2H⁺ + 2e⁻ =
Y³⁺ + 3e⁻ = Y
Zn²⁺ + 2e⁻ = Zn
Zn²⁺ + 2e⁻ = Zn(Hg)

REDUCTION

2H⁺ + 2e⁻ = H₂
Cu₂²⁺ + e⁻ = Cu +
Ge⁴⁺ + 2e⁻ = Ge²⁺
NO₃⁻ + H₂O + 2e⁻ =
Ti₂O₃ + 3H₂O +
SeO₄²⁻ + H₂O + 2e⁻ =
UO₂²⁺ + e⁻ = UO₂⁺
Pd(OH)₂ + 2e⁻ =
AgBr + e⁻ = Ag +
S₂O₈²⁻ + 2e⁻ = 2S₂O₈²⁻
AgSCN + e⁻ = Ag +
N₂ + 2H₂O + 6e⁻ =
HgO + H₂O + 2e⁻ =
Ir₂O₃ + 3H₂O +
2NO + 2e⁻ = N₂
[Co(NH₃)₆]³⁺ + e⁻ =
Hg₂O + H₂O +
Ge⁴⁺ + 4e⁻ = G
Hg₂Br₂ + 2e⁻ =
Pt(OH)₂ + 2e⁻ =
S + 2H⁺ + 2e⁻ =
Np⁴⁺ + e⁻ = Np
Ag₄[Fe(CN)₆] +
IO₃⁻ + 2H₂O +
Mn(OH)₃ + e⁻ =
2NO₂ + 3H₂O
Sn⁴⁺ + 2e⁻ =
Sb₂O₃ + 6H⁺
Cu²⁺ + e⁻ = C
BiOCl + 2H⁺
Bi(ClO₄)₃ + 3e⁻ =
Co(OH)₃ + e⁻ =
SO₄²⁻ + 4H⁺
SbO⁺ + 2H⁺
AgCl + e⁻ =
As₂O₃ + 6H⁺
Calomel electr.
Ge²⁺ + 2e⁻ =
Calomel electu
PbO₂ + H₂O
HASO₂ + 3H⁺
Ru³⁺ + e⁻ =
ReO₂ + 4H⁺
IO₃⁻ + 3H₂C
Hg₂Cl₂ + 2e⁻ =
Calomel elec
Calomel ele

Table 1 (continued)
ALPHABETICAL LISTING

E°, V	Reaction	E°, V	Reaction	E°, V
1.83	$Hg_2(ac)_2 + 2e \rightleftharpoons 2Hg + 2(ac)^-$	0.51163	$HNO_2 + H^+ + e \rightleftharpoons NO + H_2O$	0.983
0.108	$Hg_2Br_2 + 2e \rightleftharpoons 2Hg + 2Br^-$	0.13923	$2HNO_2 + 4H^+ + 4e \rightleftharpoons H_2N_2O_2 + 2H_2O$	0.86
-0.73	$Hg_2Cl_2 + 2e \rightleftharpoons 2Hg + 2Cl^-$	0.26808	$2HNO_2 + 4H^+ + 4e \rightleftharpoons N_2O + 3H_2O$	1.297
0.17	$Hg_2HPO_4 + 2e \rightleftharpoons 2Hg + HPO_4^{2-}$	0.6359	$NO_2^- + H_2O + 3e \rightleftharpoons NO + 2OH^-$	-0.46
-0.199	$Hg_2I_2 + 2e \rightleftharpoons 2Hg + 2I^-$	-0.0405	$2NO_2^- + 2H_2O + 4e \rightleftharpoons N_2^- + 4OH^-$	-0.18
-0.913	$Hg_2O + H_2O + 2e \rightleftharpoons 2Hg + 2OH^-$	0.123	$2NO_2^- + 3H_2O + 4e \rightleftharpoons N_2O + 6OH^-$	0.15
-0.407	$HgO + H_2O + 2e \rightleftharpoons Hg + 2OH^-$	0.0977	$NO_3^- + 3H^+ + 2e \rightleftharpoons HNO_2 + H_2O$	0.934
-0.744	$Hg_2SO_4 + 2e \rightleftharpoons 2Hg + SO_4^{2-}$	0.6125	$NO_3^- + 4H^+ + 3e \rightleftharpoons NO + 2H_2O$	0.957
7 H ₂ O	1.232	0.5355	$2NO_3^- + 4H^+ + 2e \rightleftharpoons N_2O_4 + 2H_2O$	0.803
-1.2	$I_2 + 2e \rightleftharpoons 2I^-$	0.536	$NO_3^- + H_2O + 2e \rightleftharpoons NO_2^- + 2OH^-$	0.01
H ₂ O	1.350	0.7	$2NO_3^- + 2H_2O + 2e \rightleftharpoons N_2O_4 + 4OH^-$	-0.85
5 OH ⁻	-0.13	1.601	$Na^+ + e \rightleftharpoons Na$	-2.71
-1.48	$H_2IO_6 + H^+ + 2e \rightleftharpoons IO_3^- + 3H_2O$	1.439	$Nb^{3+} + 3e \rightleftharpoons Nb$	-1.099
-2.92	$2HIO + 2H^+ + 2e \rightleftharpoons I_2 + 2H_2O$	0.987	$Nb_2O_5 + 10H^+ + 10e \rightleftharpoons 2Nb + 5H_2O$	-0.644
0.521	$HIO + H^+ + 2e \rightleftharpoons I^- + H_2O$	0.485	$Nd^{3+} + 3e \rightleftharpoons Nd$	-2.431
0.153	$IO^- + H_2O + 2e \rightleftharpoons I^- + 2OH^-$	1.195	$Ni^{2+} + 2e \rightleftharpoons Ni$	-0.257
0.3419	$2IO_3^- + 12H^+ + 10e \rightleftharpoons I_2 + 6H_2O$	1.085	$Ni(OH)_2 + 2e \rightleftharpoons Ni + 2OH^-$	-0.72
0.345	$IO_3^- + 6H^+ + 6e \rightleftharpoons I^- + 3H_2O$	0.15	$NiO_2 + 4H^+ + 2e \rightleftharpoons Ni^{2+} + 2H_2O$	1.678
1.103	$IO_3^- + 2H_2O + 4e \rightleftharpoons IO^- + 4OH^-$	0.26	$NiO_2 + 2H_2O + 2e \rightleftharpoons Ni(OH)_2 + 2OH^-$	-0.490
0.00	$IO_3^- + 3H_2O + 6e \rightleftharpoons I^- + 6OH^-$	-0.14	$Np^{3+} + 3e \rightleftharpoons Np$	-1.856
I^-	-0.360	-0.40	$Np^{4+} + e \rightleftharpoons Np^{3+}$	0.147
+ H ₂ O	-0.222	-0.49	$NpO_2 + H_2O + H^+ + e \rightleftharpoons Np(OH)_3$	-0.962
-0.080	$In^{3+} + e \rightleftharpoons In^{2+}$	-0.443	$O_2 + 2H^+ + 2e \rightleftharpoons H_2O_2$	0.695
-0.0034	$In^{3+} + 2e \rightleftharpoons In^+$	-0.3382	$O_2 + 4H^+ + 4e \rightleftharpoons 2H_2O$	1.229
-0.044	$In^{3+} + 3e \rightleftharpoons In$	1.156	$O_2 + H_2O + 2e \rightleftharpoons HO_2^- + OH^-$	-0.076
-3.395	$Ir^{3+} + 3e \rightleftharpoons Ir$	0.8665	$O_2 + 2H_2O + 2e \rightleftharpoons H_2O_2 + 2OH^-$	-0.146
-2.407	$[IrCl_4]^- + e \rightleftharpoons [IrCl_4]^{2-}$	0.77	$O_2 + 2H_2O + 4e \rightleftharpoons 4OH^-$	0.401
-0.36	$[IrCl_4]^{2-} + 3e \rightleftharpoons Ir + 6Cl^-$	0.098	$O_3 + 2H^+ + 2e \rightleftharpoons O_2 + H_2O$	2.076
3.053	$IrO_3 + 3H_2O + 6e \rightleftharpoons 2Ir + 6OH^-$	-2.931	$O_3 + H_2O + 2e \rightleftharpoons O_2 + 2OH^-$	1.24
2.866	$K^+ + e \rightleftharpoons K$	-2.522	$O(g) + 2H^+ + 2e \rightleftharpoons H_2O$	2.421
2.153	$La^{3+} + 3e \rightleftharpoons La$	-2.90	$OH + e \rightleftharpoons OH^-$	2.02
-0.447	$La(OH)_3 + 3e \rightleftharpoons La + 3OH^-$	-3.0401	$HO_2^- + H_2O + 2e \rightleftharpoons 3OH^-$	0.878
-0.037	$Li^+ + e \rightleftharpoons Li$	-2.70	$OsO_4 + 8H^+ + 8e \rightleftharpoons Os + 4H_2O$	0.85
0.771	$Mg^{2+} + 2e \rightleftharpoons Mg$	-2.372	$P(red) + 3H^+ + 3e \rightleftharpoons PH_3(g)$	-0.111
0.358	$Mg(OH)_2 + 2e \rightleftharpoons Mg + 2OH^-$	-2.690	$P(white) + 3H^+ + 3e \rightleftharpoons PH_3(g)$	-0.063
H ₂ O	2.20	-1.185	$P + 3H_2O + 3e \rightleftharpoons PH_3(g) + 3OH^-$	-0.87
-0.56	$Mn^{2+} + 2e \rightleftharpoons Mn$	1.5415	$H_2PO_2^- + e \rightleftharpoons P + 2OH^-$	-1.82
phen) ₃] ²⁺	1.147	1.224	$H_3PO_2 + H^+ + 3e \rightleftharpoons P + 2H_2O$	-0.508
1 mol/l	1.00	0.558	$H_3PO_3 + 2H^+ + 2e \rightleftharpoons H_3PO_2 + H_2O$	-0.499
0.400	$MnO_2 + 4H^+ + 3e \rightleftharpoons Mn^{2+} + 2H_2O$	1.679	$H_3PO_3 + 3H^+ + 3e \rightleftharpoons P + 3H_2O$	-0.454
-0.560	$MnO_2 + 8H^+ + 5e \rightleftharpoons Mn^{2+} + 4H_2O$	1.507	$HPO_3^{2-} + 2H_2O + 2e \rightleftharpoons H_2PO_2^- + 3OH^-$	-1.65
OH ⁻	-1.21	0.595	$HPO_3^{2-} + 2H_2O + 3e \rightleftharpoons P + 5OH^-$	-1.71
0.24	$MnO_2 + 2H_2O + 3e \rightleftharpoons MnO_2 + 4OH^-$	0.60	$H_3PO_4 + 2H^+ + 2e \rightleftharpoons H_3PO_3 + H_2O$	-0.276
0.12	$MnO_2 + 2H_2O + 2e \rightleftharpoons MnO_2 + 4OH^-$	-1.56	$PO_4^{3-} + 2H_2O + 2e \rightleftharpoons HPO_3^{2-} + 3OH^-$	-1.05
0.00	$Mn(OH)_2 + 2e \rightleftharpoons Mn + 2OH^-$	0.15	$Pb^{2+} + 2e \rightleftharpoons Pb$	-0.1262
O	-0.118	-0.200	$Pb^{2+} + 2e \rightleftharpoons Pb(Hg)$	-0.1205
H ₂ O	-0.18	0.092	$PbBr_2 + 2e \rightleftharpoons Pb + 2Br^-$	-0.284
0.000	$N_2 + 2H_2O + 6H^+ + 6e \rightleftharpoons 2NH_4OH$	-3.09	$PbCl_2 + 2e \rightleftharpoons Pb + 2Cl^-$	-0.2675
-2.23	$N_2 + 2H^+ + 2e \rightleftharpoons 2NH_4^+$	1.275	$PbF_2 + 2e \rightleftharpoons Pb + 2F^-$	-0.3444
-1.99	$N_2 + 3H^+ + 2e \rightleftharpoons 2NH_4^+$	1.766	$PbHPO_4 + 2e \rightleftharpoons Pb + HPO_4^{2-}$	-0.465
-0.33	$N_2O + 2H^+ + 2e \rightleftharpoons N_2 + H_2O$	2.65	$PhI_2 + 2e \rightleftharpoons PhI + 2I^-$	-0.365
-1.776	$N_2O + 2H^+ + 2e \rightleftharpoons N_2 + 2H_2O$	0.867	$PbO + H_2O + 2e \rightleftharpoons Pb + 2OH^-$	-0.580
O	-1.75	1.065	$PbO_2 + 4H^+ + 2e \rightleftharpoons Pb^{2+} + 2H_2O$	1.455
O	-1.25	1.035	$HPO_3^{2-} + H_2O + 2e \rightleftharpoons Pb + 3OH^-$	-0.537
4 OH ⁻	-0.21	1.42	$PbO_2 + H_2O + 2e \rightleftharpoons PbO + 2OH^-$	0.247
0.00	$NH_4OH + H^+ + 2e \rightleftharpoons N_2H_5^+ + 2H_2O$	0.10	$PbO_2 + SO_4^{2-} + 4H^+ + 2e \rightleftharpoons PbSO_4 + 2H_2O$	1.6913
0.00	$NO + 2H^+ + 2e \rightleftharpoons N_2O + H_2O$	1.591	$PbSO_4 + 2e \rightleftharpoons Pb + SO_4^{2-}$	-0.3588
0.00	$NO + H_2O + 2e \rightleftharpoons N_2O + 2OH^-$	0.76		

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Table 1 (continued)
ALPHABETICAL LISTING

Reaction	E°, V	Reaction	E°, V
p-benzoquinone + 2 H ⁺ + 2 e ⁻ ⇌ hydroquinone	0.6992	Co ³⁺ + e ⁻ ⇌ Co ²⁺ (2 mol/l H ₂ SO ₄)	1.8
BiCl ₄ ⁻ + 3 e ⁻ ⇌ Bi ⁻ + 4 Cl ⁻	0.16	[Co(NH ₃) ₆] ³⁺ + e ⁻ ⇌ [Co(NH ₃) ₆] ²⁺	0.108
-Bi ₂ O ₃ + 3 H ₂ O + 6 e ⁻ ⇌ 2 Bi + 6 OH ⁻	-0.46	Co(OH) ₂ + 2 e ⁻ ⇌ Co + 2 OH ⁻	0.16
-Bi ₂ O ₃ + 4 H ⁺ + 2 e ⁻ ⇌ 2 BiO ⁺ + 2 H ₂ O	-1.593	Co(OH) ₃ + e ⁻ ⇌ Co(OH) ₂ + OH ⁻	0.16
-BiO ⁺ + 2 H ⁺ + 3 e ⁻ ⇌ Bi + H ₂ O	0.320	CO ₂ + 2 H ⁺ + 2 e ⁻ ⇌ HCOOH	0.13
-BiOCl + 2 H ⁺ + 3 e ⁻ ⇌ Bi + Cl ⁻ + H ₂ O	0.1583	Cr ³⁺ + 2 e ⁻ ⇌ Cr	0.01
-Br ₂ (aq) + 2 e ⁻ ⇌ 2 Br ⁻	1.0873	Cr ³⁺ + e ⁻ ⇌ Cr ²⁺	-0.407
Br ₂ (l) + 2 e ⁻ ⇌ 2 Br ⁻	1.066	Cr ³⁺ + 3 e ⁻ ⇌ Cr	-0.74
HBrO + H ⁺ + 2 e ⁻ ⇌ Br ⁻ + H ₂ O	1.331	Cr ₂ O ₇ ²⁻ + 14 H ⁺ + 6 e ⁻ ⇌ 2 Cr ³⁺ + 7 H ₂ O	1.23
HBrO + H ⁺ + e ⁻ ⇌ 1/2 Br ₂ (aq) + H ₂ O	1.574	CrO ₂ + 2 H ₂ O + 3 e ⁻ ⇌ Cr + 4 OH ⁻	-1.2
HBrO ₂ + H ⁺ + e ⁻ ⇌ 1/2 Br ₂ (l) + H ₂ O	1.596	HCrO ₄ ⁻ + 7 H ⁺ + 3 e ⁻ ⇌ Cr ³⁺ + 4 H ₂ O	-1.350
BrO ⁻ + H ₂ O + 2 e ⁻ ⇌ Br ⁻ + 2 OH ⁻	0.761	CrO ₄ ²⁻ + 4 H ₂ O + 3 e ⁻ ⇌ Cr(OH) ₃ + 5 OH ⁻	-0.13
BrO ₃ ⁻ + 6 H ⁺ + 5 e ⁻ ⇌ 1/2 Br ₂ + 3 H ₂ O	1.482	Cr(OH) ₃ + 3 e ⁻ ⇌ Cr + 3 OH ⁻	-1.4
BrO ₃ ⁻ + 6 H ⁺ + 6 e ⁻ ⇌ Br ⁻ + 3 H ₂ O	1.423	Cs ⁺ + e ⁻ ⇌ Cs	-2.92
BrO ₃ ⁻ + 3 H ₂ O + 6 e ⁻ ⇌ Br ⁻ + 6 OH ⁻	0.61	Cu ⁺ + e ⁻ ⇌ Cu	-0.52
Ca ⁺ + e ⁻ ⇌ Ca	-3.80	Cu ²⁺ + e ⁻ ⇌ Cu ⁺	0.15
Ca ²⁺ + 2 e ⁻ ⇌ Ca	-2.868	Cu ²⁺ + 2 e ⁻ ⇌ Cu	-0.345
Calomel electrode, 1 molal KCl	0.2800	Cu ²⁺ + 2 e ⁻ ⇌ Cu(Hg)	-1.103
Calomel electrode, 1 mol/l KCl (NCE)	0.2801	Cu ²⁺ + 2 CN ⁻ + e ⁻ ⇌ [Cu(CN) ₂] ⁻	0.00
Calomel electrode, 0.1 mol/l KCl	0.3337	CuI ₂ + e ⁻ ⇌ Cu + 2 I ⁻	-0.360
Calomel electrode, saturated KCl (SCE)	0.2412	Cu ₂ O + H ₂ O + 2 e ⁻ ⇌ 2 Cu + 2 OH ⁻	-0.222
Calomel electrode, saturated NaCl (SSCE)	0.2360	Cu(OH) ₂ + 2 e ⁻ ⇌ Cu + 2 OH ⁻	-0.080
Ca(OH) ₂ + 2 e ⁻ ⇌ Ca + 2 OH ⁻	-3.02	2 Cu(OH) ₂ + 2 e ⁻ ⇌ Cu ₂ O + 2 OH ⁻ + H ₂ O	-0.003
Cd ²⁺ + 2 e ⁻ ⇌ Cd	-0.4030	D ⁺ + e ⁻ ⇌ 1/2 D ₂	-0.044
Cd ²⁺ + 2 e ⁻ ⇌ Cd(Hg)	-0.3521	2 D ⁺ + 2 e ⁻ ⇌ D ₂	-3.395
Cd(OH) ₂ + 2 e ⁻ ⇌ Cd(Hg) + 2 OH ⁻	0.809	Eu ²⁺ + 2 e ⁻ ⇌ Eu	-2.407
CdSO ₄ + 2 e ⁻ ⇌ Cd + SO ₄ ²⁻	-0.246	Eu ³⁺ + 3 e ⁻ ⇌ Eu	-0.36
Ce ³⁺ + 3 e ⁻ ⇌ Ce	-2.483	Eu ³⁺ + e ⁻ ⇌ Eu ²⁺	-0.36
Ce ³⁺ + 3 e ⁻ ⇌ Ce(Hg)	-1.4373	F ₂ + 2 H ⁺ + 2 e ⁻ ⇌ 2 HF	3.053
Ce ⁴⁺ + e ⁻ ⇌ Ce ³⁺	1.61	F ₂ + 2 e ⁻ ⇌ 2 F ⁻	2.866
CeOH ³⁺ + H ⁺ + e ⁻ ⇌ Ce ³⁺ + H ₂ O	1.715	F ₂ O + 2 H ⁺ + 4 e ⁻ ⇌ H ₂ O + 2 F ⁻	-2.153
Cl ₂ (g) + 2 e ⁻ ⇌ Cl ⁻	1.35827	Fe ²⁺ + 2 e ⁻ ⇌ Fe	-0.447
HClO + H ⁺ + e ⁻ ⇌ 1/2 Cl ₂ + H ₂ O	1.611	Fe ³⁺ + 3 e ⁻ ⇌ Fe	-0.037
HClO + H ⁺ + 2 e ⁻ ⇌ Cl ⁻ + H ₂ O	1.482	Fe ³⁺ + e ⁻ ⇌ Fe ²⁺	0.771
ClO ⁻ + H ₂ O + 2 e ⁻ ⇌ Cl ⁻ + 2 OH ⁻	0.81	[Fe(CN) ₆] ³⁻ + e ⁻ ⇌ [Fe(CN) ₆] ⁴⁻	0.358
ClO ₂ + H ⁺ + e ⁻ ⇌ HClO ₂	1.277	FeO ₄ ²⁻ + 8 H ⁺ + 3 e ⁻ ⇌ Fe ³⁺ + 4 H ₂ O	2.20
HClO ₂ + 2 H ⁺ + 2 e ⁻ ⇌ HClO + H ₂ O	1.645	Fe(OH) ₃ + e ⁻ ⇌ Fe(OH) ₂ + OH ⁻	-0.56
HClO ₂ + 3 H ⁺ + 3 e ⁻ ⇌ 1/2 Cl ₂ + 2 H ₂ O	1.628	[Fe(phenanthroline) ₃] ³⁺ + e ⁻ ⇌ [Fe(phen) ₃] ²⁺	1.147
HClO ₂ + 3 H ⁺ + 4 e ⁻ ⇌ Cl ⁻ + 2 H ₂ O	1.570	[Fe(phen) ₃] ³⁺ + e ⁻ ⇌ [Fe(phen) ₃] ²⁺ (1 mol/l H ₂ SO ₄)	1.06
ClO ₂ ⁻ + H ₂ O + 2 e ⁻ ⇌ ClO ⁻ + 2 OH ⁻	0.66	[Ferricinium] ⁺ + e ⁻ ⇌ ferrocene	0.400
ClO ₂ ⁻ + 2 H ₂ O + 4 e ⁻ ⇌ Cl ⁻ + 4 OH ⁻	0.76	Ga ³⁺ + 3 e ⁻ ⇌ Ga	-0.560
ClO ₂ (aq) + e ⁻ ⇌ ClO ₂ ⁻	0.954	H ₂ GaO ₃ ⁻ + H ₂ O + 3 e ⁻ ⇌ Ga + 4 OH ⁻	-1.219
ClO ₃ ⁻ + 2 H ⁺ + e ⁻ ⇌ ClO ₂ + H ₂ O	1.152	Ge ²⁺ + 2 e ⁻ ⇌ Ge	0.24
ClO ₃ ⁻ + 3 H ⁺ + 2 e ⁻ ⇌ HClO ₂ + H ₂ O	1.214	Ge ⁴⁺ + 4 e ⁻ ⇌ Ge	0.124
ClO ₃ ⁻ + 6 H ⁺ + 5 e ⁻ ⇌ 1/2 Cl ₂ + 3 H ₂ O	1.47	Ge ⁴⁺ + 2 e ⁻ ⇌ Ge ²⁺	0.00
ClO ₃ ⁻ + 6 H ⁺ + 6 e ⁻ ⇌ Cl ⁻ + 3 H ₂ O	1.451	GeO ₂ + 2 H ⁺ + 2 e ⁻ ⇌ GeO + H ₂ O	-0.118
ClO ₃ ⁻ + H ₂ O + 2 e ⁻ ⇌ ClO ₂ ⁻ + 2 OH ⁻	0.33	H ₂ GeO ₃ + 4 H ⁺ + 4 e ⁻ ⇌ Ge + 3 H ₂ O	-0.182
ClO ₃ ⁻ + 3 H ₂ O + 6 e ⁻ ⇌ Cl ⁻ + 6 OH ⁻	0.62	2 H ⁺ + 2 e ⁻ ⇌ H ₂	0.00000
ClO ₄ ⁻ + 2 H ⁺ + 2 e ⁻ ⇌ ClO ₃ ⁻ + H ₂ O	1.189	H ₂ + 2 e ⁻ ⇌ 2 H ⁻	-2.23
ClO ₄ ⁻ + 8 H ⁺ + 7 e ⁻ ⇌ 1/2 Cl ₂ + 4 H ₂ O	1.39	HO ₂ + H ⁺ + e ⁻ ⇌ H ₂ O ₂	1.495
ClO ₄ ⁻ + 8 H ⁺ + 8 e ⁻ ⇌ Cl ⁻ + 4 H ₂ O	1.389	2 H ₂ O + 2 e ⁻ ⇌ H ₂ + 2 OH ⁻	-0.8277
ClO ₄ ⁻ + H ₂ O + 2 e ⁻ ⇌ ClO ₃ ⁻ + 2 OH ⁻	0.36	H ₂ O ₂ + 2 H ⁺ + 2 e ⁻ ⇌ 2 H ₂ O	1.776
(CN) ₂ + 2 H ⁺ + 2 e ⁻ ⇌ 2 HCN	0.373	HfO ₂ ⁺ + 2 H ⁺ + 4 e ⁻ ⇌ Hf + H ₂ O	-1.724
2 HCNO + 2 H ⁺ + 2 e ⁻ ⇌ (CN) ₂ + 2 H ₂ O	0.330	HfO ₂ + 4 H ⁺ + 4 e ⁻ ⇌ Hf + 2 H ₂ O	-1.503
(CNS) ₂ + 2 e ⁻ ⇌ 2 CNS ⁻	0.77	HfO(OH) ₂ + H ₂ O + 4 e ⁻ ⇌ Hf + 4 OH ⁻	-2.50
Co ³⁺ + 2 e ⁻ ⇌ Co	-0.28	Hg ²⁺ + 2 e ⁻ ⇌ Hg	0.851
		2 Hg ²⁺ + 2 e ⁻ ⇌ Hg ₂ ²⁺	0.920
		Hg ₂ ²⁺ + 2 e ⁻ ⇌ 2 Hg	0.7973

A = Weight per cent of solute
B = Molecular weight of solvent
E = Molecular weight of solute
F = Grams of solute per liter of solution

G = Molality
M = Molarity
N = Mole fraction
R = Density of solution grams per cc

Concentration of solute— SOUGHT	Concentration of solute—GIVEN				
	A	N	G	M	F
A	—	$\frac{100N \times E}{N \times E + (1 - N)B}$	$\frac{100G \times E}{1000 + G \times E}$	$\frac{M \times E}{10R}$	$\frac{F}{10R}$
N	$\frac{\frac{A}{E} + \frac{100 - A}{B}}{E}$	—	$\frac{B \times G}{B \times G + 1000}$	$\frac{B \times M}{M(B - E) + 1000R}$	$\frac{B \times F}{F(B - E) + 1000R \times E}$
G	$\frac{1000A}{E(100 - A)}$	$\frac{1000N}{B - N \times B}$	—	$\frac{1000M}{1000R - (M \times E)}$	$\frac{1000F}{E(1000R - F)}$
M	$\frac{10R \times A}{E}$	$\frac{1000R \times N}{N \times E + (1 - N)B}$	$\frac{1000R \times G}{1000 + E \times G}$	—	$\frac{F}{E}$
F	10AR	$\frac{1000R \times N \times E}{N \times E + (1 - N)B}$	$\frac{1000R \times G \times E}{1000 + G \times E}$	M × E	—

ELECTROCHEMICAL SERIES

Petr Vanýsek

There are three tables for this Electrochemical Series. Each table lists standard reduction potentials, E° values, at 298.15 K (25°C), and at a pressure of 101.325 kPa (1 atm.). Table 1 is an alphabetical listing of the elements according to the symbols for the elements. Thus, data for Silver (Ag) precedes those for Aluminum (Al). Table 2 lists only those reduction reactions which have E° values positive to the potential of the Standard Hydrogen Electrode. In Table 2, the reactions are listed in the order of increasing positive potential and range from 0.000 V to +3.053 V. Table 3 lists only those reduction reactions which have E° values negative to the potential of the Standard Hydrogen Electrode. In Table 3, reactions are listed in the order of increasing negative potential and range from -0.017 to -4.10 V.

Table 1
ALPHABETICAL LISTING

Reaction	E° , V	Reaction	E° , V
$\text{Ag}^+ + e \rightleftharpoons \text{Ag}$	0.7996	$\text{Ag}_2\text{WO}_4 + 2e \rightleftharpoons 2\text{Ag} + \text{WO}_4^{2-}$	0.4660
$\text{Ag}^{2+} + e \rightleftharpoons \text{Ag}^+$	1.980	$\text{Al}^{3+} + 3e \rightleftharpoons \text{Al}$	-1.662
$\text{Ag}(\text{ac}) + e \rightleftharpoons \text{Ag} + (\text{ac})^-$	0.643	$\text{H}_2\text{AlO}_3^- + \text{H}_2\text{O} + 3e \rightleftharpoons \text{Al} + 4\text{OH}^-$	-2.33
$\text{AgBr} + e \rightleftharpoons \text{Ag} + \text{Br}^-$	0.07133	$\text{AlF}_6^{3-} + 3e \rightleftharpoons \text{Al} + 6\text{F}^-$	-2.069
$\text{AgBrO}_3 + e \rightleftharpoons \text{Ag} + \text{BrO}_3^-$	0.546	$\text{As} + 3\text{H}^+ + 3e \rightleftharpoons \text{AsH}_3$	-0.608
$\text{Ag}_2\text{C}_2\text{O}_4 + 2e \rightleftharpoons 2\text{Ag} + \text{C}_2\text{O}_4^{2-}$	0.4647	$\text{As}_2\text{O}_3 + 6\text{H}^+ + 6e \rightleftharpoons 2\text{As} + 3\text{H}_2\text{O}$	0.234
$\text{AgCl} + e \rightleftharpoons \text{Ag} + \text{Cl}^-$	0.22233	$\text{HAsO}_2 + 3\text{H}^+ + 3e \rightleftharpoons \text{As} + 2\text{H}_2\text{O}$	0.248
$\text{AgCN} + e \rightleftharpoons \text{Ag} + \text{CN}^-$	-0.017	$\text{AsO}_2^- + 2\text{H}_2\text{O} + 3e \rightleftharpoons \text{As} + 4\text{OH}^-$	-0.68
$\text{Ag}_2\text{CO}_3 + 2e \rightleftharpoons 2\text{Ag} + \text{CO}_3^{2-}$	0.47	$\text{H}_3\text{AsO}_4 + 2\text{H}^+ + 2e \rightleftharpoons \text{HAsO}_2 + 2\text{H}_2\text{O}$	0.560
$\text{Ag}_2\text{CrO}_4 + 2e \rightleftharpoons 2\text{Ag} + \text{CrO}_4^{2-}$	0.4470	$\text{AsO}_4^{3-} + 2\text{H}_2\text{O} + 2e \rightleftharpoons \text{AsO}_2^- + 4\text{OH}^-$	-0.71
$\text{AgF} + e \rightleftharpoons \text{Ag} + \text{F}^-$	0.779	$\text{Au}^+ + e \rightleftharpoons \text{Au}$	1.692
$\text{Ag}[\text{Fe}(\text{CN})_6] + 4e \rightleftharpoons 4\text{Ag} + [\text{Fe}(\text{CN})_6]^{4-}$	0.1478	$\text{Au}^{3+} + 2e \rightleftharpoons \text{Au}^+$	1.401
$\text{AgI} + e \rightleftharpoons \text{Ag} + \text{I}^-$	-0.15224	$\text{Au}^{3+} + 3e \rightleftharpoons \text{Au}$	1.498
$\text{AgIO}_3 + e \rightleftharpoons \text{Ag} + \text{IO}_3^-$	0.354	$\text{AuBr}_2^+ + e \rightleftharpoons \text{Au} + 2\text{Br}^-$	0.959
$\text{Ag}_2\text{MoO}_4 + 2e \rightleftharpoons 2\text{Ag} + \text{MoO}_4^{2-}$	0.4573	$\text{AuBr}_4^- + 3e \rightleftharpoons \text{Au} + 4\text{Br}^-$	0.854
$\text{AgNO}_3 + e \rightleftharpoons \text{Ag} + \text{NO}_3^-$	0.564	$\text{AuCl}_4^- + 3e \rightleftharpoons \text{Au} + 4\text{Cl}^-$	1.002
$\text{Ag}_2\text{O} + \text{H}_2\text{O} + 2e \rightleftharpoons 2\text{Ag} + 2\text{OH}^-$	0.342	$\text{Au}(\text{OH})_3 + 3\text{H}^+ + 3e \rightleftharpoons \text{Au} + 3\text{H}_2\text{O}$	1.45
$\text{Ag}_2\text{O}_2 + \text{H}_2\text{O} + 2e \rightleftharpoons 2\text{Ag} + 2\text{OH}^-$	0.739	$\text{H}_2\text{BO}_3^- + 5\text{H}_2\text{O} + 8e \rightleftharpoons \text{BH}_4^- + 8\text{OH}^-$	1.24
$2\text{Ag}_2\text{O} + \text{H}_2\text{O} + 2e \rightleftharpoons 2\text{Ag}_2\text{O} + 2\text{OH}^-$	0.607	$\text{H}_2\text{BO}_3^- + \text{H}_2\text{O} + 3e \rightleftharpoons \text{B} + 4\text{OH}^-$	-1.79
$\text{AgOCN} + e \rightleftharpoons \text{Ag} + \text{OCN}^-$	0.41	$\text{H}_3\text{BO}_3 + 3\text{H}^+ + 3e \rightleftharpoons \text{B} + 3\text{H}_2\text{O}$	0.8698
$\text{Ag}_2\text{S} + 2e \rightleftharpoons 2\text{Ag} + \text{S}^{2-}$	-0.691	$\text{Ba}^{2+} + 2e \rightleftharpoons \text{Ba}$	-2.912
$\text{Ag}_2\text{S} + 2\text{H}^+ + 2e \rightleftharpoons 2\text{Ag} + \text{H}_2\text{S}$	-0.0366	$\text{Ba}^{2+} + 2e \rightleftharpoons \text{Ba}(\text{Hg})$	-1.570
$\text{AgSCN} + e \rightleftharpoons \text{Ag} + \text{SCN}^-$	0.08951	$\text{Ba}(\text{OH})_2 + 2e \rightleftharpoons \text{Ba} + 2\text{OH}^-$	2.99
$\text{Ag}_2\text{SeO}_3 + 2e \rightleftharpoons 2\text{Ag} + \text{SeO}_3^{2-}$	-0.3629	$\text{Be}^{2+} + 2e \rightleftharpoons \text{Be}$	1.847
$\text{Ag}_2\text{SO}_4 + 2e \rightleftharpoons 2\text{Ag} + \text{SO}_4^{2-}$	-0.654	$\text{Be}_2\text{O}_3 + 3\text{H}_2\text{O} + 4e \rightleftharpoons 2\text{Be} + 6\text{OH}^-$	-2.63